

**GRANDE PRAIRIE REGIONAL COLLEGE
ACADEMIC UPGRADING DEPARTMENT**

**CHEMISTRY 0120
COURSE OUTLINE**

INSTRUCTOR: NANCY FRASER

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OFFICE HOURS: As posted on my office door. Other times are available by appointment.

PREREQUISITES: CH 0110

COREQUISITES: MA 0110

COURSE GOALS: This course is designed to provide the student with an understanding of the following chemical concepts: Bonding, Chemical Equations, Stoichiometry, Solutions, and Organic Chemistry.

TEXT BOOK: Chemistry: A Study of Matter, by Dorin, Demmin, Gabel,
4th edition, 1992.

**ADDITION
COSTS:** Lab coat

EVALUATION: Regular attendance is expected of all students, and is crucial to passing the course. Students who miss classes will soon find themselves falling behind and failing. Lateness will **not** be tolerated as it interrupts the instructor and fellow classmates. As per Department Policy, if you miss more than 15 days per semester (approximately 1 day/week) in any course, you may be debarred from the final exam for that course.

A certificate (a doctor's or a note from the funeral home) will be required to make up the final exam. **You will receive a grade of F if you miss the final.** Call if you are going to miss a test. There may be a deduction of 10% for test rewrites.

*****Very important:** **Laboratory attendance to each specific experiment is compulsory; a passing grade in the laboratory component is required to pass the course.** There are NO 'make up' labs in this course.

Lab reports must be submitted on the required date and at the **required time.**

Assignments will not be accepted after the assignment has been returned to the class. I am usually a speedy marker and return papers the next day.

Penalties for late **assignments** are as follows: (Assuming that I have not returned the marked assignments

1 day late – 20%, 1 days late – 50%, 3 days late – 100%

Penalties for late **lab reports** are as follows:

5 minutes – 10%, 24 hours – 20%, after that – 100%

Marking Scheme:

Lab Reports:	15%
Assignments:	15%
Tests:	15%
Midterm:	15%
Final Exam:	<u>40%</u>
Total	100%

PLAGIARISM AND CHEATING POLICY:

See College Calendar.

GRADING EQUIVALENCE THAT WILL BE USED IN THIS COURSE:

Alpha Grade	4-point Equivalent	Percentage Guidelines	Designation
A ⁺	4.0	90 – 100	EXCELLENT
A	4.0	85 – 89	
A ⁻	3.7	80 – 84	FIRST CLASS STANDING
B ⁺	3.3	76 – 79	
B	3.0	73 – 75	GOOD
B ⁻	2.7	70 – 72	
C ⁺	2.3	67 – 69	SATISFACTORY
C	2.0	64 – 66	
C ⁻	1.7	60 – 63	
D ⁺	1.3	55 – 59	MINIMAL PASS
D	1.0	50 – 54	
F	0	0 – 49	FAIL

Course Content

Unit I	REVIEW	Pages
a.	Metric review	16 – 24
b.	Significant Figures	26 – 56
c.	Definition of : chemistry, matter, density, states of matter, classification of matter, properties of matter	63 – 79
d.	Elements and their symbols.	
e.	Nomenclature*	147 – 166

*** Nomenclature is one of the most important topics that you will learn at the secondary level. It will NOT be review at the post secondary level. If you are having trouble with this topic get help IMMEDIATELY!!! See me!**

Unit II ORGANIC CHEMISTRY (Chapter 24)

It maybe necessary to move this until after stoichiometry.

On completing this section, you should be able to:

- a. Define organic compounds and organic chemistry.
- b. Name the first ten alkanes using IUPAC System. Write the general formula of alkanes.
- c. Explain structural formula and isomerism.
- d. Name the first ten alkenes and alkynes and write their general formulas.
- e. Define functional groups. Identify the functional groups of halides, alcohols, aldehydes, ketones, amines and carboxylic acids.
- f. Write structural formula(s) for a compound with given molecular formula.

Unit III Biochemistry (Chapter 25) (As time permits)

Will always follow organic chemistry.

On completing this section, you should understand (at a basic level):

- a. Carbohydrates
- b. Lipids
- c. Proteins
- d. Biochemical reactions and enzymes
- e. Nucleic acids

Unit IV **STOICHIOMETRY (Chapter 8, 10, 16)**

On completing this section, you should be able to:

- Define formula(s) and molecular masses.
- Define the mole and the Avogadro's number.
- Convert masses of a substance in g into moles and vice-versa.
- Determine the percentage composition of a compound if its formula is known.
- Determine the empirical and the molecular formulas of compounds if their percentage compositions are known.
- Define chemical equation.
- Write chemical equations for simple reactions.
- Balance a given chemical equation.
- Interpret a chemical equation in terms of the moles of various reactants and the products.
- Given the balanced equation for a reaction and the number of moles of one of the reactants or a product, be able to calculate the number of moles of all other species.
- Given the mass of one of the reactants and the products, be able to calculate the masses or the moles of all the other species.
- Perform calculations based on masses, volumes of gases or molarities of solutions.

*** MIDTERM

Unit V: **ATOMIC STRUCTURE (Chapter 6 & pages 77 – 79)**

(Some old stuff and a lot of new stuff)

(There will be some copies of Basic chemistry in the Learning Center and on reserve in the library)

On completing this section, you should be able to:

- Define the subatomic particles - electrons, protons and neutrons - and give their relative masses and charges.
- Describe Dalton's Atomic Theory and explain how it has been modernized.
- Describe Thomson's model of the atom and its shortcomings.
- Describe Rutherford model of the atom and explain how it was improved by Bohr.
- Express the relation between protons, neutrons, atomic number and mass number.
- Explain isotopes and average atomic masses.
- Solve problems involving the per cents of various isotopes of an element and its atomic mass.
- Define principal energy levels and give their maximum capacity to hold electrons.
- Distribute electrons in principal energy levels.
- Define orbitals and explain different types of orbitals.
- Distribute electrons in various orbitals.
- Understand the basics of emission and absorption spectroscopy.
- Be able to explain the importance of spectroscopy in identifying substances.

Unit VI CHEMICAL BONDING (Chapter15)

On completing this section, you should be able to:

- a. Write Lewis dot structures for elements if their atomic numbers are given.
- b. Define chemical bond and valence electrons.
- c. Explain ionic and covalent bonds giving examples of each.
- d. Predict the nature of bond(s) between two elements from their Lewis dot structures or atomic numbers.
- e. Explain double and triple bonds.
- f. Discuss the nature of bonding between elements of common compounds and ions.
- g. Discuss the shapes of molecules based on VSEPR theory.
- h. Explain electronegativity and the polarity of molecules.
- i. Explain hydrogen bonding, metallic bonding, and intermolecular forces.

Unit VII PERIODIC TABLE (Chapter 14)

On completing this section, you should be able to:

- a. State the modern periodic law.
- b. Explain groups and periods.
- c. Relate the electronic configuration of elements to groups and periods.
- d. Locate the various groups of elements on the periodic table.
- e. Describe the characteristics of representative and transition elements.
- f. Discuss in detail the periodic trends, such as atomic size, metallic character, ionization energy and electronegativity.

Unit VIII PHASES OF MATTER (CHAPTER 11)

On completing this section, you should be able to:

- a. Define boiling point, evaporation, and condensation
- b. Compare and explain the characteristic properties of gases, liquids, and solids using the kinetic molecular theory.
- c. Relate vapour pressure to boiling point.
- d. Explain how bonding types and molecular mass can affect boiling points and vapour pressure.
- e. Account for the unusual behavior of water.

Unit XI SOLUTIONS (Chapter 16)

On completing this section, you should be able to:

- a. Define solute, solvent and solution.
- b. Discuss different types of solutions, such as gas, liquid, or solid solutions.
- c. Define solubility and describe various factors that affect solubility.
- d. Define saturated, unsaturated and supersaturated solutions.
- e. Define concentration of solutions and give the units of molarity.
- f. Work out problems involving moles, volumes and molarity.